“Isotopic” perspective on the Periodic Table of Elements, following Dr. Peter Plichta

While we are long accustomed to looking at the Table, there is more information to it that we routinely overlook. With Dr. Plichta’s insights we can mine this information (see also Plichta.de)

First While ‘atomic number’ might have served a purpose while the Table was being compiled, what useful information does it now convey? What purpose does it now serve?

So I substitute the term “Proton Count” instead. It was always true, if not appreciated. The useful information conveyed is the Element’s identity. It also distinguishes from the Neutron Count, which is also important, though for different reasons. It also distinguishes from the Nucleon Count, which is also important. The Neutron and the Nucleon count vary, but if the Proton Count does not, we continue to deal with the same Element.

Second While ‘atomic weight’ might have served a purpose while the Table was being compiled, what useful information does it now convey? What useful purpose does it now serve?

What we see routinely represented in the Table, and taught, is a combination of isotopic admixture and measurement error. How can that serve a useful purpose? Indeed, it instead obscures valuable information, that is, the number of Isotopes of each Element, and the distribution of those numbers of Isotopes across the Elements and the Periods of the Periodic Table of Elements.

Does it surprise that there appears to be a pattern. To be sure, we sometimes see patterns where there is no underlying phenomenon, but that examination is required to determine that. When has that examination occurred? How can any determination yet be made?

Dr. Plichta drew attention to the Single (also Pure) and to the Double Isotopes. My attempt to examine the Triple, and then the Multiple (4 or more) Isotopes served little purpose in another tabular presentation, so I resorted to this one. That is when the pattern seemed to emerge. I have color coded it so as to ease distinguishing among the various Elemental Isotopes. RadioActive (non-stable) Elements and Isotopes (Tritium, etc.) are distinguished yet again.

Exactly that examination is requested. We can argue about any determination afterward.

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Table 1a

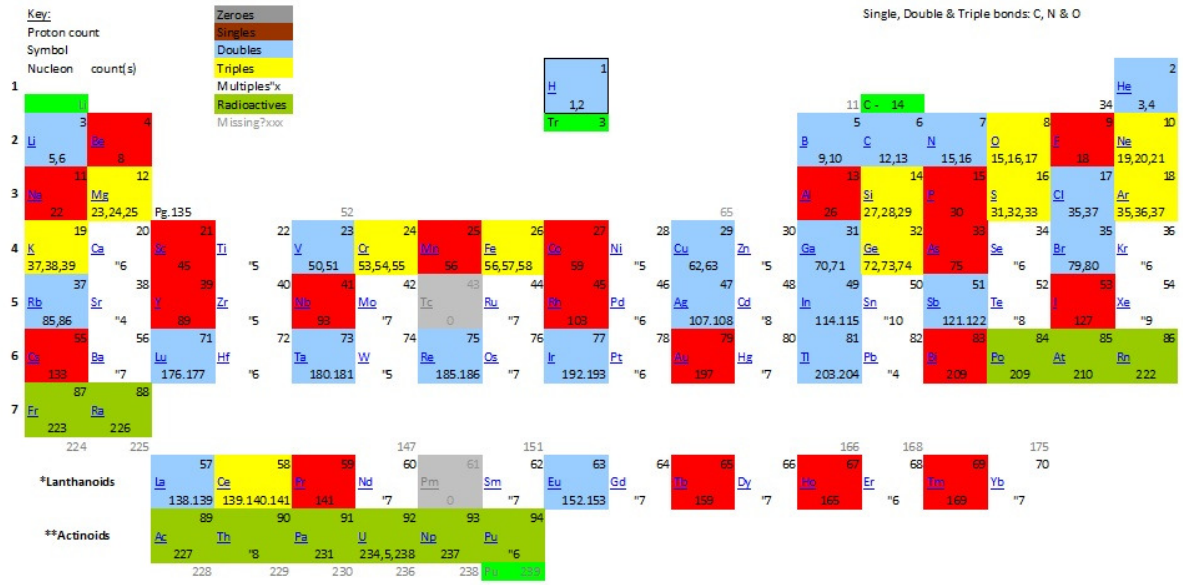

Table 1b

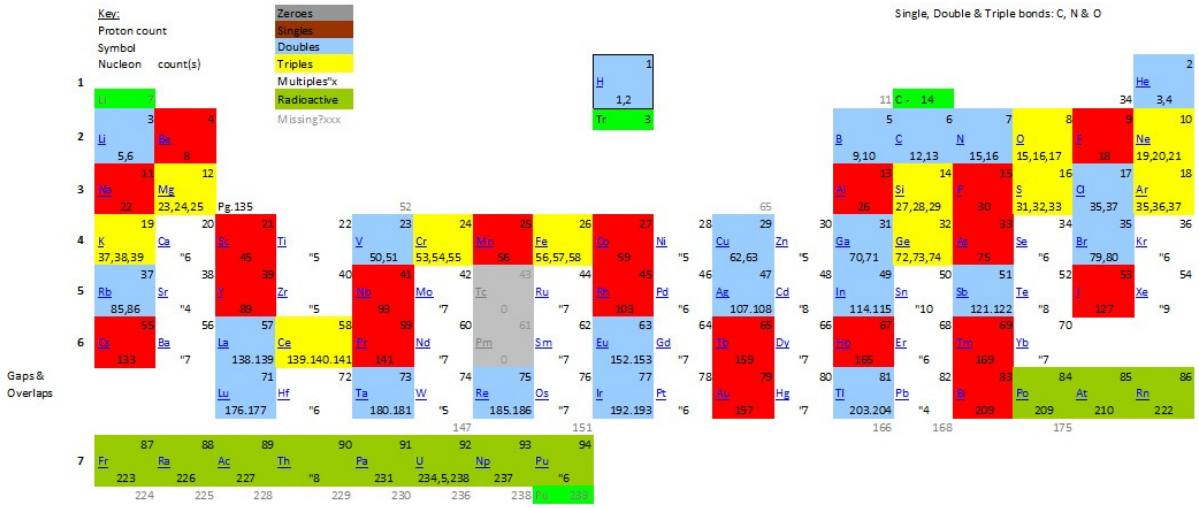

Table 2

From Dr. Peter Plichta's book "God's Secret Formula" English translation, Element books, mid-90's, page 134

count	Pure (single) Isotopes	Additional Neutrons	Double Isotopes	Additional Neutrons	Triple Isotopes	Additional Neutrons
1	Be - 4		He - 2		O - 8	
2	F - 9		Li - 3		Ne - 10	
3	Na - 11		B - 5		Mg - 12	
4	Al - 13		N - 7		Si - 14	
5	P - 15		Cl - 17		S - 16	
6	Sc - 21	3	V - 23	4, 5	Ar - 18	
7	Mn - 25	6	Cu - 29	4, 5	K - 19	
8	Co - 27	5	Ga - 31	8, 9	Cr - 24	5, 6, 7
9	As - 33	9	Br - 35	9, 10	Fe - 26	4, 5, 6
10	Y - 39	11	Rb - 37	11, 12	Ge - 32	8, 9, 10
11	Nb - 41	11	Ag - 47	13, 14	Ce - 58	22, 23, 24
12	Rh - 45	13	In - 49	16, 17		
13	I - 53	21	Sb - 51	18, 20		
14	Cs - 55	23	La - 57	24, 25		
15	Pr - 59	23	Eu - 63	26, 27		
16	Tb - 65	29	Lu - 71	34, 35		
17	Ho - 67	31	Ta - 73	34, 35		
18	Tm - 69	31	Re - 75	35, 36		
19	Au - 79	39	Ir - 77	38, 39		
20	Bi - 83	43	Tl - 81	41, 42		

Literature

Dr. Peter Plichta, polymath, "God's Secret Formula" Element books, 1997
also Plichta.de

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Dr. Georg Michlo, Ph.D. "The Push of Gravity" Vantage books, mid-90's

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"Isotopes in Chemistry" by Duncan & Cock